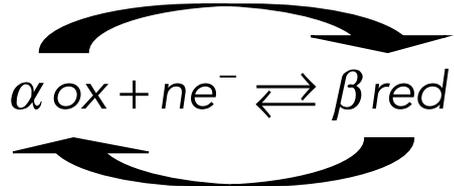


Réduction = arrivée d'électrons, n.o ↓

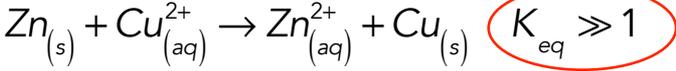
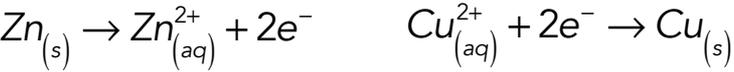


Oxydation = départ d'électrons, n.o ↑

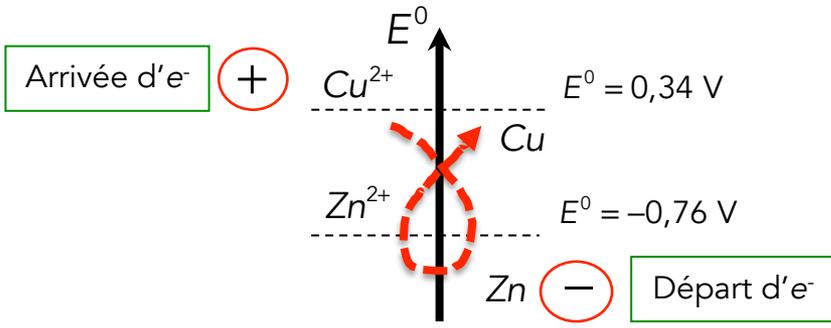
RELATION DE NERNST

$$E_{(ox/red)} = E_{(ox/red)}^0 + \frac{RT}{nF} \ln \left(\frac{a(\text{côté ox})^\alpha}{a(\text{côté red})^\beta} \right)$$

Oxydation=ANODE Réduction=CATHODE



$$\Delta E^0 = \frac{RT}{nF} \ln(K_{eq}) \leftarrow E_{cathode} = E_{anode} \quad \text{EQUILIBRE}$$



- Rappels:**
- n.o(H) = +I
 - n.o(O) = -II
- sauf dans H₂O₂ n.o(O) = -I
et dans O₂ n.o(O) = 0